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**Chapter 7 Ionic
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The energy required to separate 1 mol of ions in an ionic compound is referred to as lattice energy. Lattice energy is directly related to the size of the ions that are bonded. Smaller ions form compounds with more closely spaced ionic charges

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and require more energy to separate. Smaller, harder to get between them.

**Chemistry chapter 7:
Ionic Compounds
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206 Chapter 7 • Ionic
Compounds and Metals
Section 77.1.1 Figure
7.1 As carbon dioxide
dis-solves in ocean
water, carbonate ions
are produced. Coral
polyps capture these
car-bonate ions,

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producing crystals of calcium carbonate, which they secrete as an exo-skeleton. Over time, the coral reef forms. A coral reef is a complex habitat that

Chapter 7: Ionic Compounds and Metals - Mr. Miller

Formula Unit - the chemical formula for an ionic compound represent the simplest ratio of the ions involved. $MgCl_2$ -

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magnesium and chloride ions exist in a 1:2 ratio. Contains one Mg^{2+} ion and two Cl^- ions for a total charge of zero. ... Chapter 7: Ionic Compounds and Metals

Chapter 7: Ionic Compounds and Metals

Chapter 7: Ionic compounds names and formulas. STUDY. PLAY. Ion. An atom with a + or - charge. Cation. Ion

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with a + charge. Anion.
Ion with - charge.
Polyatomic ion. A
group of atoms that act
as a unit & carry a
charge. Binary
compound. Containing
atoms of 2 different
elements. Ternary
compounds.

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Terms in this set (19)

Ion. An atom that
gained or lost electrons
and has a charge.

Cation-Lost electron-

Positive charge-Metals.

Anion-Gained electrons-

Negative charge-

Nonmetals. Properties.

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Chemistry: Matter and
Change • Chapter 7

103 Section 7.1 Ion
Formation pages

206–209 Section 7.1

Assessment page 209

1. Compare the
stability of a lithium
atom with that of its
ion, Li^+ . The Li^+ ion is
more stable because it

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has a complete octet.

2.

**Ionic Compounds
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Compound Naming
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oxidation numbers and
correct subscript latex
codes)

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Chapter 7- Ionic Compounds and Metals - Chemistry with ...

Naming Ions. The name of a monatomic cation is simply the name of the element followed by the word ion. Thus, Na^+ is the sodium ion, Al^{3+} is

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Answers

the aluminum ion, Ca^{2+} is the calcium ion, and so forth.. We have seen that some elements lose different numbers of electrons, producing ions of different charges (Figure 3.3).

5.7: Naming Ionic Compounds - Chemistry LibreTexts

A compound that contains ions and is held together by ionic

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bonds is called an ionic compound. The periodic table can help us recognize many of the compounds that are ionic: When a metal is combined with one or more nonmetals, the compound is usually ionic.

3.7 Molecular and Ionic Compounds - General Chemistry 1 & 2

A compound that

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contains ions and is held together by ionic bonds is called an ionic compound. The periodic table can help us recognize many of the compounds that are ionic: When a metal is combined with one or more nonmetals, the compound is usually ionic.

2.6 Molecular and Ionic Compounds - Chemistry

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In every ionic compound, the total number of positive charges of the cations equals the total number of negative charges of the anions. Thus, ionic compounds are electrically neutral overall, even though they contain positive and negative ions. We can use this observation to help us write the formula of an ionic compound.

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Chapter 7 Chem:

Practice Test . 48

Questions | By Konlee |

Last updated: Jan 9, ...

The crystal lattice of
ionic compounds affect
their melting and

boiling points. A. True.

B. False. 41. The lattice

energy is the energy
required to separate
the ions of an ionic

compound. A. True. B.

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7: Ionic and Metallic
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Ionic solids are also poor conductors of electricity for the same reason—the strength of ionic bonds prevents ions from moving freely in the solid state. Most ionic solids, however, dissolve readily in water. Once dissolved or melted, ionic compounds are excellent conductors of electricity and heat because the ions can

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move about freely.

**Ionic Bonding |
General Chemistry -
Lecture & Lab**

Classification of
compounds:

compounds are of
following types. 1)

Ionic compounds:

Examples: NaCl, KBr,
CuSO₄. ... 9th Class

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