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Study Guide 1. Each family in the periodic table has its own characteristic properties based upon its number of _____. 2. The valence electrons are those electrons _____ from the nucleus. 3. In an electron dot diagram of carbon, _____ dots should be drawn around the

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element's symbol.

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Based on the types of atoms bonding (metals and/or nonmetals), classify if the bond is an ionic bond or a covalent bond for:
CaO=ionic,
HF=covalent, CO₂=
covalent, and
LiCl=ionic. What is a diatomic molecule?

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Molecule containing
only two atoms

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Unit 3- Chemical
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CHEMICAL BONDS 1.

How are atoms held
together? What is the
attraction between? 2.

What are the three
characteristics of
atoms that contribute
to bonding? Explain

each. 3. What is a
solute? What is a
solvent? Give an

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example of each. 4. For each of the 4 different types of bonds, explain the following a. Bond ...

Unit 3- Chemical Bonding Study Guide

Chemical Bonding Study Guide 01) 3.

According to the octet rule, chemical compounds tend to form so that each atom has an octet of electrons in its highest energy level.

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**Honors Chemistry -
Home**

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_____. Write the
electron configuration
for each atom below
and circle which ones
are stable. Fe b. P³⁻ c.
Ar d. Cs¹⁺.

[Ar]4s²3d⁶-not stable
[Ne]3s²3p⁶-stable [Ne]
3s²3p⁶-stable
[Xe]6s²-not stable.

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what is a compound? a

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pure substance
composed of two or
more elements that are
chemically combined.
name 3 familiar
compounds. vinegar,
water, table salt.

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A double bond between
two carbon atoms.
Notice that each
carbon achieves eight
electrons by this

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sharing. Because each shared pair constitutes a single covalent bond, the two shared pairs are called a double bond. The structure on the right side of Figure 8 shows this double bond of four shared electrons with two lines, and the left side of Figure

Covalent Bonds - CliffsNotes Study Guides

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4) - Study Guide 1 Ions made up of more than one atom are called —
A polar compound B nonpolar compound C polyatomic ions D metallic bonds
2 How many electrons do all halogen family elements need to gain in order to attain a noble gas configuration?

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When these strategies
are employed, the
atoms become linked
by a chemical bond.

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Bonds are connections
between atoms. A solid
grasp of valence shell
electron pair repulsion
(VSEPR) theory will
help you understand

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how elements that differ by one or two atomic numbers behave. According to VSEPR theory, the number of electrons an element has corresponds with its chemical properties.

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Compounds with
Additional Elements
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Quiz: Ionic Bonds -
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Julina Nocera Chemical
Bonding Unit

Objectives: Study
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Differentiate between
intra- and inter-
molecular bonds/forces
Intermolecular forces
occur between

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molecules,
intramolecular bonds
occur between atoms
2. Identify metallic
bonds and how they
account for the
properties of metals 3.

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A chemical bond is a
force of attraction
between atoms. It
occurs when atoms
share or transfer
electrons. A chemical

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compound is a new substance that forms when atoms of different elements form chemical bonds. A compound always consists of a fixed ratio of elements.

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Chemical Bonds Study
Guide 11 November
2016 A molecule with
covalent bond o
Formula unit with ionic

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bond * Molecules:
formed by covalent
bonds * Lattice energy
is the energy released
in the formation of an
ionic compound. The
formation of an IONIC
BOND is the result of
the transfer of one or
more electrons from a
metal onto a non-
metal.

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Characteristics of both

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bonds: * Occur between 2 atoms * Composed of 2 electrons * Have both ionic and covalent characteristics * Together = 100% * Both bonds are measured on an electronegativity scale * Both contain a nonmetal * Chemical bonds * Are determined by using the "magic number" (1.67) * Have bond angle and bond axis

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ecf8427e.