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126 Chemistry: Matter and Change • Chapter 8 Solutions Manual Section 8.5 Electronegativity and Polarity pages 265-270 Section 8.5 Assessment page 270 68. Summarize how electronegativity difference is related to bond character. The greater the electronegativity difference, the greater the ionic nature of the bond. 69.

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Chapter 8 - Covalent Bonding - 8 Assessment - Page 257: 68 Answer Boiling points require the molecules to separate , when the state changes from a liquid into a gas.

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210 Chapter 8 Ionic Compounds CHAPTER 8 What You'll Learn You will define a chemical bond. You will describe how ions form. You will identify ionic bonding and the characteristics of ionic compounds. You will name and write formulas for ionic compounds. You will relate metallic bonds to the characteristics of metals. Why It's Important

Chapter 8: Ionic Compounds

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Section 8.1 Assessment page 247 7. Identify the type of atom that generally forms covalent bonds. The majority of covalent bonds form between nonmetallic elements. 8. Describe how the octet rule applies to covalent bonds. Atoms share valence electrons; the shared electrons complete the octet

of each atom. 9. Illustrate the formation of single, double, and

Covalent Bonding

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Mastering chemistry Chapter 8. An attractive interaction between a hydrogen atom bonded to a very electronegative atom (O, N, or F). In one molecule and an electron rich regions in a different molecule. (Also could appear within a single molecule) It commonly bonds to unshared electron pairs on another electronegative atom.

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