

Iron Nail In An Aqueous Solution Answer

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Iron Nail In An Aqueous

Step 1 of 5. (a) When iron nail is placed in copper sulphate (CuSO_4) solution, iron is reacted with CuSO_4 . forms ferrous sulphate (FeSO_4) solution because iron is more reactive than copper so it replace the copper from CuSO_4 . The chemical reaction is as follows:

Solved: When a clean iron nail is placed in an aqueous ...

Write two observations that you will make when an iron nail is kept in an aqueous solution of copper sulphate. Write the chemical equation for this reaction. 0 votes . 1.8k views. asked Oct 29, 2017 in Class X Science by akansha Expert (2.2k points)

Write two observations that you will make when an iron ...

(a) Chemical reaction of iron nail with copper sulphate solution in water. THEORY Iron displaces copper ions from an aqueous solution of copper sulphate. It is a single displacement reaction of one metal by another metal. Iron is placed above copper in the activity series. Elements placed above in this series are more reactive than those placed below them.

Materials - NCERT

Chem Help Please!!When a clean iron nail is placed in an aqueous solution of copper(II) sulfate, the nail imme? (a) What is the material coating the iron? (b) What is the oxidizing agent? (I thought CuSO_4 , but its incorrect.) (c) Write the balanced equation for the reaction. (Use the lowest possible coefficients.

Chem Help Please!!When a clean iron nail is placed in an ...

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Iron Nail In An Aqueous Solution Answer

If a nail made of elemental Iron [$\text{Fe}(s)$] is placed in an aqueous solution of a soluble palladium(II) salt [$\text{Pd}^{2+}(aq)$] the nail will gradually disappear as the iron enters the solution $\text{Fe}^{3+}(aq)$ and...

Show the net ionic equation? | Yahoo Answers

Consult the EMF activity series and predict which of the following changes will occur if an iron nail is placed in an aqueous solution containing the (Sn^{2+}) ion. A) Iron will be oxidized. B) Tin will be further oxidized. C) Nothing at all will happen (no reaction). D) Iron will be reduced.

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In this case, iron, being more reactive than copper, will displace it from the solution. That is, ferrous ions or Fe (II) will go into the solution, while curic or Cu (II) will be deposited onto...

What happens when an iron nail is immersed in copper ...

When an iron nail is dipped in copper sulphate solution, a brown coating of copper is formed on the surface of iron and the colour of copper sulphate solution changes from blue to light green. This reaction shows that iron is more reactive than copper as it displaces copper from its solution and iron passes into solution as Fe

When Iron nails are added to an aquous solution of copper ...

Consult the EMF activity series and predict which of the following changes will occur if an iron nail is placed in an aqueous solution containing the (Sn^{2+}) ion. Iron will be oxidized.

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A third technique applies to situations (such as an automobile radiator) where aqueous solutions are in contact with the iron. Corrosion inhibitors include chromate salts and organic compounds such as tribntylamine, ($\text{C}_4\text{H}_9\text{N}$)₃. Chromates apparently form an impervious coating of $\text{FeCrO}_4(s)$ as soon as any iron is oxidized to iron(II).

22.7: Corrosion - Chemistry LibreTexts

Expert Answer: (a) When an iron nail is immersed in the aqueous solution of copper sulphate, a brown coating of copper is developed over iron and the blue copper sulphate solution becomes pale green since iron is more reactive than copper. (b) $\text{Fe}(s) + \text{CuSO}_4(aq) \rightarrow \text{FeSO}_4(aq) + \text{Cu}(s)$

what happened when an iron nail is kept immersed in copper ...

The iron can be solid or aqueous but the copper sulphate must be aqueous in order to facilitate the reaction. There are several physical indicators of the reaction between iron and CuSO_4 . Aqueous CuSO_4 is a blue solution, and it will lose its color as it reacts with the iron to form FeSO_4 . Once the reaction is complete and there is no more CuSO_4 present in the solution, it will appear completely colorless.

What Happens When Fe and CuSO_4 React?

An iron nail was dipped in a salt solution. After sometime a reddish brown deposition of the nail was seen. The salt solution could be (a) Silver nitrate (b) Sodium sulphate (c) Aluminium chloride (d) Copper sulphate. 24. The colour of the residue left in the test tube after heating ferrous sulphate which undergoes decomposition is: (a) yellowish-brown

NCERT Class 10 Science Lab Manual Types of Reactions ...

Solution for If 27.0 g of iron nail was used to react with oxygen gas (O_2). Calculate the mass of ferric oxide produced in this reaction