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the rate of reaction (a)
1 g of sugar (cubes) 1
gram of sugar (grains)
Y surface area (b) 50°C
0°C X temperature (c)
low number of particles
= few collisions high
number of particles =
more collisions Y
concentration (d)
enzyme added no
enzyme added X
catalyst (e) twigs logs
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Section 6.2 Factors Affecting the Rate of Chemical Reactions

Chapter 18 Reaction
Rates and

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graph below shows the
reaction rate of the
same reaction with and
without a catalyst. Use
it to help you answer
Questions 18 and 19.
18. Label each curve
as with catalyst or
without catalyst. 19.

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What does the graph show about the effect of a catalyst on the rate of a reaction? 20.

Name Date Class **REACTION RATES** **AND EQUILIBRIUM** **18**

Determine the integrated rate law, the differential rate law, and the value of the rate constant.

Calculate the $[H_2O_2]$ at 4000. s after the start of the reaction. A

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